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= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, [HC 2H 3O 2(aq)], can be determined: 23 2 23 2 HC H O (aq) 23 2 HC H O (aq) 23 2 [HC H O (aq)] (6.60 mmol) = 38.00 mL

### Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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