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= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, [HC₂H₃O₂(aq)], can be determined: $\frac{6.60 \text{ mmol}}{38.00 \text{ mL}} = 0.174 \text{ mol/L}$

Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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