

## Oxidation Reduction Concept Review Answers

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### Oxidation Reduction Concept Review Answers

Oxidation Reduction Concept Review Answers Oxidation is the loss of electrons or the addition of oxygen; reduction is the gain of electrons or the addition of hydrogen.  $Al \rightarrow Al^{3+} + 3e^{-}$  (oxidation);  $O_2 + 2e^{-} \rightarrow 2O^{2-}$  (reduction) (answers will vary) 5.5: Oxidation-Reduction (Redox) Reactions - Chemistry ... Answer. Reduction:  $Ca^{2+} + 2e^{-} \rightarrow Ca$ .

### Oxidation Reduction Concept Review Answers

Question: Oxidation-Reduction Reactions Concept Review 8.77. How Are The Gains Or Losses Of Electrons Related To Changes In Oxidation Numbers? 8.78. What Is The Sum Of The Oxidation Numbers Of The Atoms In A Molecule? 8.79.

### Oxidation-Reduction Reactions Concept Review 8.77 ...

Answer. Reduction:  $Ca^{2+} + 2e^{-} \rightarrow Ca$ . Oxidation:  $2(K \rightarrow K^{+} + e^{-})$  Combined:  $Ca^{2+} + 2K \rightarrow Ca + 2K$  The substance oxidized is the reactant that had undergone oxidation: K; The substance reduced is the reactant that had undergone reduction:  $Ca^{2+}$  The reducing agent is the same as the substance oxidized: K

### 6.1: Oxidation-Reduction (Redox) Reactions - Chemistry ...

Oxidation is the loss of electrons or the addition of oxygen; reduction is the gain of electrons or the addition of hydrogen.  $Al \rightarrow Al^{3+} + 3e^{-}$  (oxidation);  $O_2 + 2e^{-} \rightarrow 2O^{2-}$  (reduction) (answers will vary)

### 5.5: Oxidation-Reduction (Redox) Reactions - Chemistry ...

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### The role of oxidation states to decide whether an ...

The sum of the oxidation numbers of the elements in a neutral compound is 0. In a polyatomic ion, however, the sum is equal to the charge. Oxidation numbers help you keep track of electron transfer in redox reactions. An oxidation number increase is oxidation, while a decrease is reduction.

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### The oxygen-reduction reaction needs to be identified ...

Oxidation reaction: the atom or ion becomes more positive due to the loss of negatively charged electrons. Reduction: an atom or ion gains one or more electrons ( $e^{-}$ ): reactants. product. Reduction reaction:  $X + e^{-} \rightarrow X^{-}$ . Reduction: electrons ( $e^{-}$ ) are a reactant in the reduction reaction.

### Oxidation and Reduction Concepts Chemistry Tutorial

Questions pertaining to redox reactions. If you're seeing this message, it means we're having trouble loading external resources on our website.

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